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many moles of
hydrogen atoms
are in 1 mole of
each of the

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following

compounds? H 2

0 _____ 2 _____ CH 3

OH _____ 4 _____ (NH

4) 2 CO 3

_____ 8 _____ 2.

Empirical and

Molecular

Formulas Give

the empirical

formula for each

compound: N 2 H

2 _____ NH _____ BH 3

_____ BH 3 _____ H

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202 OH For
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The atomic
ratios in each
compound are
also the
relative number
of atomic mass
units of its

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elements. The

first example is
nitrous oxide (N_2O), as shown

in Table 1. The
relative masses
were obtained by
multiplying the
atomic ratios
and atomic
masses. ...

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In the reaction

represented by

the equation 2Na

$2\text{H}_2\text{O} \rightarrow 2\text{NaOH} +$

H_2 , how many

grams of

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hydrogen are
produced if 120.
g of Na and 80.0
g 12 1
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the quantitative
relationships

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stoichiometry?

1. true
2. true
3. false
4. true
5. true
6. 2, 2,
- 64.10
7. 3, 3,
- 96.00
8. 2, 2,

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88.02 9. 4, 4,

72.08 10.

methanol and

oxygen gas 11.

carbon dioxide

and water 12.

160.10 g 13.

160.10 g 14.

They are equal.

15. A mole ratio

is a ratio

between ...

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