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*Unit 7 Endothermic*

*reaction/graph* **Chemistry,**

**class-12, unit-7** Chemistry

~~Review Sheet Unit 7~~

are representative of each type of problem worked in this unit. For more practice on a given type of problem, refer back to Note Taking Guides, worksheets, and quizzes.

1. Calculate the molar mass for ethane, C<sub>2</sub>H<sub>6</sub>.  
C<sub>2</sub>H<sub>6</sub> 2(12.0 g) + 6(1.0 g) = 30.0 g

2. Convert 45 g Se to moles of Se. ?  
mol Se = 45 g Se ! 1 mol Se 79.0 g Se = 0.57 mol Se

3.

~~Review Sheet: Unit 7 Name~~

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~~Vocabulary: Fill in the blanks ...~~  
are representative of each type of problem worked in this unit. For more practice on a given type of problem, refer back to Note Taking Guides, worksheets, and quizzes.

1. Calculate the molar mass for ethane,  $C_2H_6$ .
2. Convert 45 g Se to moles of Se.
3. Convert  $4.77 \times 10^{24}$  molecules of  $SO_2$  to grams.

~~Review Sheet: Unit 7 Name~~

~~Vocabulary: Fill in the blanks ...~~

UNIT 7 REVIEW SHEET -

EQUILIBRIUM • EQUILIBRIUM is the point at which the rate of the forward reaction is equal to the rate of the reverse reaction in a reversible process. o Reaction is dynamic - still occurring - even though to the eye, it looks like

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nothing is changing at equilibrium. o Observable/measurable quantities (concentration, pressure, color) are constant at equilibrium. o The amounts of products and reactants are typically not equal at equilibrium!

~~UNIT 7 REVIEW SHEET.pdf~~ ~~UNIT 7 REVIEW SHEET \u2013 2013 ...~~

CHEMISTRY CP Name: \_\_\_\_\_ Per: \_\_\_\_\_ TEST: \_\_\_\_\_ Unit 7 Review Sheet: Thermochemistry KEY  
Potential Energy Diagrams:  
Choose the appropriate letter. 1. What is the  $\Delta H$  of the forward reaction? f 2. What is the activation energy of the forward reaction? b 3. What is the H of the reverse reaction? f 4.

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Review Sheet: Unit 5 Name \_\_\_\_\_

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2004, GPB 5.24a KEY I. For each of the following statements, write "I" for ionic, "C" for covalent

~~CHEMISTRY: A Study of Matter~~

Review Sheet for Chemistry Unit 5 December 11, 2011 31. Balance the following equations by putting the proper coefficients wherever they are needed. a)  $\text{Sr} + \text{O}_2 \rightarrow \text{SrO}$   
b)  $\text{Al} + \text{H}_2\text{O} \rightarrow \text{Al}(\text{OH})_3 + \text{H}_2$  c)  $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$  d)  $\text{C}_5\text{H}_{12} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$  e)  $\text{Na} + \text{H}_2\text{O} \rightarrow \text{H}_2 + \text{NaOH}$   
f)  $\text{NaClO}_3 \rightarrow \text{NaCl} + \text{O}_2$  g) ...

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